HOMEWORK TWO

1.A. Complete the following statements :

1.The number of elements which exist in nature is .....

2. Elements of s-block are located on the……they arranged in..... groups.

3. The period number of the element equals the number of ..... occupied by electrons in its atom.

4. The transition elements start from period……..

5. The modern periodic table consists of…… horizontal periods and ……vertical groups.

B. What is the scientific principle upon which the elements are arranged in the modern periodic table ?

2.A. Write the scientific term:

1. Ablock of elements which is located on the right side of the periodic table.

2.A kind of elements that are located in d-block.

3.The block which contains the groups from (3A) to (0).

4.It indicates the number of electrons that revolve in energy levels.

5. The block which contains the lanthanides and actinides.

B. Locate the position of the following elements in the modern periodic table:

1. $_{11}Na$:

2. $_{18}Ar$:

3. $_{20}Ca$:

4. $_{9}F$:

C. From the following table :

	Groups							
	1A	2A	3A	4A	5A	6A	7A	0
Second period					Y			
Third period	X							Z

1. Calculate the atomic number of element (X).

2. What is the modern number of the group of the element (Y) ?

3. Complete : The outermost energy level of element (Z) contains

3. A. Choose the correct answer :

1. Elements of p-block are arranged in .... group.

a. two

b. five

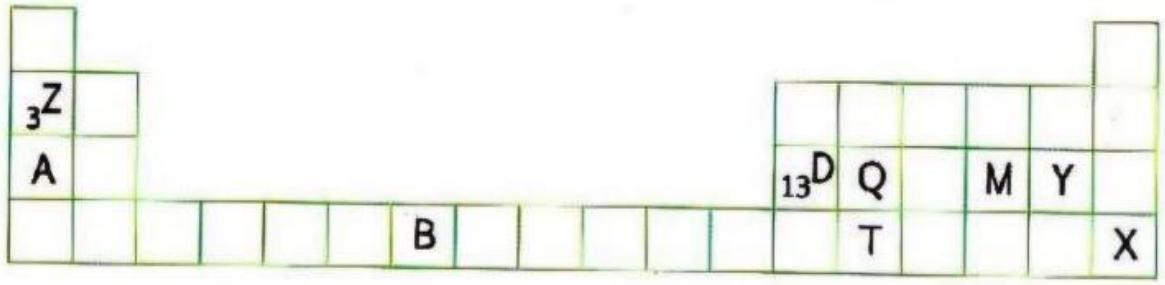
c. six

d. eight

2. The element that its atomic number equals 17 is similar in its chemical properties to the element that its atomic number equals……

a.2 b.9 c.7 d.11

B. The following table represents a section of the modern periodic table:



1.What is the symbol indicates each of the following:

a. A transition element :

b. An inert gas :

c. An element lies in the third period and group (6A):

2. What is thc block of the elements (A), (B) and (D) ?

3. Choose: The atomic number of element (B) is…….element (T).

a. more than

b. equal to

c. Less than

4.A. Find the atomic number of each of the following:

1. An element exists in period 2 and group (6A).
2. An element exists in period 3 and group (1A).
3. An element exists in period 2 and group (0).

B. Give a reason for:

Elements of the same group have similar properties.